<u>Chemistry 3610</u>	Inorganic Chemistry		
Lecturer	Adam S. Veige: veige@chem <u>CLB 412b</u>	.ufl.edu	392-9844
	Office Hours: M and Th (10:50 CLB 414) am – 12:00 pm)	
<u>Teaching Assistants</u>	Matias Pascualini: <u>matiaspascu</u> Sudarsan VenkatRamani: <u>sud.v</u> (Lab: CLB 415)		· · · · · · · · · · · · · · · · · · ·
	Office Hours: To Be Determin	ed	
Lecture Hours	M, T, W, R, F Period 2 (9:30 – 10:45)		
<u>Textbook</u>	Miessler, G. L. and Tarr, D. A., Inorganic Chemistry 5th Ed.		
<u>Helpful Text</u>	Shriver and Atkins Inorganic Chemistry 5th Ed. Cotton, Wilkinson, Gauss, Advanced Inorganic Chemistry Cotton, Chemical Applications of Group Theory		
Grading		200	
	Mid-term Exams (2)	200	
	Problem Sets 6	60 150	
	<u>Final Exam</u> Total	<u>150</u> 410	
	1.0001		

To review the current UF grade point equivalencies go to: https://catalog.ufl.edu/ugrad/current/regulations/info/grades.aspx#averaging

	· · · · ·	328-314 B+, 233-218 D+,		296-287 B-, 201-187 D-,	-
		or problem set.	.** After two hotocopy exar	weeks the score ns and problem	
<u>Examinatio</u>	<u>ns</u>			ay 29 th , Friday, . ¹ , in class 9:30 -	
<u>Missed Exa</u>	<u>ms</u>	made for stude Notification an	ents that have ond documentat	official UF trave ion must be pro	angements will be el conflicts. vided one week ee due to medical

	illness must be accompanied by a doctor note indicating you were not able to attend the exam.	
<u>Accommodation for Stud</u>	lents with Disabilities Students requesting classroom accommodation must first register with the Dean of Students Office. That office will provide the student with documentation that he/she must provide to the course instructor when requesting accommodation.	
<u>Lecture</u>	Chemistry 3610 will survey modern inorganic/organometallic concepts of bonding, reactivity, and physical properties.	
<u>Problem Sets</u>	Problem sets will be assigned at intervals of approximately one week. Problem Sets are due at the beginning of class. Problem sets handed in immediately after class but on the same day will be assigned a grade of M (5 pts). Problem sets handed in after the due date will not be graded (0 pts) Solutions will be provided.	
	Grading: Problem sets will be graded as follows Satisfactory: S (10 pts) Marginal: M (5 pts) Unsatisfactory: U (0 pts)	
	Satisfactory (S) problems were attempted and there is an obvious understanding of the material demonstrated. (i.e. just attempting a question is not satisfactory) Marginal (M) grade will be assigned for sloppy work, not attempting a problem, if a significant portion is incorrect. Unsatisfactory (U) majority of the problem sets is incorrect.	
<u>Problem Set Due dates</u>	*Subject to change Friday, May 15 Friday, May 22 Thursday, May 28 Friday, June 5 Thursday, June 11 Thursday, June 18	
<u>Review Sessions</u>	Each Friday the lecture will be a review session. The topic will be the problem set/exam answer key and to answer general questions.	
<u>Class Attendance</u>	Class attendance is mandatory since some discussion may diverge from the text.	

<u>Honesty Policy</u>	All students admitted to the University of Florida have signed a statement of academic honesty committing themselves to be honest in all academic work and understanding that failure to comply with this commitment will result in disciplinary action. This statement is a reminder to uphold your obligation as a UF student and to be honest in all work submitted and exams taken in this course and all others.
<u>UF Counseling Services</u>	 Resources are available on-campus for students having personal problems or lacking clear career and academic goals. The resources include: University Counseling Center, 301 Peabody Hall, 392-1575, Personal and Career Counseling. SHCC mental Health, Student Health Care Center, 392-1171, Personal and Counseling. Center for Sexual Assault/Abuse Recovery and Education (CARE), Student Health Care Center, 392-1161, sexual assault counseling. Career Resource Center, Reitz Union, 392-1601, career development assistance and counseling.

Chemistry 3610: Inorganic Chemistry Course Information

Chapter 1. Atomic Structure
Chapter 2. Molecular Structure and Bonding
Chapter 3. The structure of Simple Solids
Chapter 4. Acids and Bases
Chapter 6. Molecular Symmetry
Chapter 7. An introduction to coordination compounds
Chapter 20. d-Metal complexes: electronic structure and properties
Chapter 21. Coordination chemistry: reactions of complexes.
Chapter 22. d-metal organometallic chemistry