

CHM2045 — General Chemistry I — Fall 2018

Credits: 3; Prereq: MAC 1147 or the equivalent, and a passing score on the chemistry placement exam or a passing grade in CHM 1025; Coreq: CHM 2045L.

The first semester of the CHM 2045/2045L and CHM 2046/2046L sequence. Stoichiometry, atomic and molecular structure, the states of matter, reaction rates and equilibria. A minimum grade of C is required to progress to CHM2046. A minimum grade of B+ is required to progress to CHM2051 (Honors General Chemistry II). (P)

Instructor	Dr. Alexander Angerhofer (Dr. A)
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E-mail	alex@chem.ufl.edu
O.H.	T-9, W-8, R-8, and by appointment, JHH202.

TAs	Erik Ferenczy, erikferenczy@ufl.edu , O.H.: MWF-6 (CLC*), and by appointm. Jason Smith, jason14thbrklyn@ufl.edu , O.H.: MW-7 (CLC*), and by appointm. Sam Cockey, csamuel11@ufl.edu , O.H.: T-3, F-7 (CLC*), and by appointm. Austin Chen, au.chen@ufl.edu , O.H.: M-8, T-4 (CLC*), and by appointm.	
	Sam Cockey	Jason Smith
Sections	Class #21895, Section #3F05, R-3	Class #21896, Section #3F11, R-4
E-mail	csamuel11@ufl.edu	jason14thbrklyn@ufl.edu
O.H.	CLC*	CLC*

*Chemistry Learning Center, 1st floor of JHH, room #105.

Class Meeting Times	MWF-3 periods, 9:35am–10:25am in Joseph Hernandez Hall 221	
Discussion Sessions	Class #21895, Section #3F05	Class #21896, Section #3F11
	R-3 in MAT 119	R-4 in LEI 207
Holidays	09/03 (Labor Day), 11/02 (Homecoming), 11/12 (Veterans Day), 11/21–23 (Thanksgiving Holidays), 12/6–7 (Reading Days, no classes).	
Class Text	M. Silberberg, “Chemistry: The Molecular Nature of Matter and Change With Advanced Topics,” 8th Edition, McGraw-Hill, New York 2018, ISBN: 978-1259741098. The text is recommended. Any reasonably recent General Chemistry textbook should be fine for you to review and look up material.	
Homework	Homework will be assigned weekly except during weeks of during-term exams,. Homework will be graded.	
Points Earnable	4 progress exams @ 15% each, for 60% total. 1 cumulative final exam @ 23%. 10 homeworks @ 1% each, for 10% total. Daily in-class participation grade (learning catalytics) @ 5% total. 1 ALEKS prep, due online Sept. 10 @ 2% total. Total earnable points are 100%.	
Grading Scheme ¹	A: $\geq 90.0\%$ 90.0% > A- $\geq 86.0\%$ 86.0% > B+ $\geq 83.0\%$ 83.0% > B $\geq 80.0\%$ 80.0% > B- $\geq 77.0\%$ 77.0% > C+ $\geq 73.0\%$ 73.0% > C $\geq 69.0\%$ 69.0% > D+ $\geq 66.0\%$ 66.0% > D $\geq 63.0\%$ 63.0% > D- $\geq 60.0\%$ 60.0% > E.	

¹ see <https://catalog.ufl.edu/ugrad/current/regulations/info/grades.aspx> for more info on UF grade policies.

Course Schedule (tentative):

Date	Day	Chap.	Topic	Reading
08/22/18	W	1	Disc. of Syllabus and review of chap. 1, Units, Unit Conversion, Greek Prefixes, Significant Figures	Chap. 1
08/23/18	R	2	Review of chap. 2, compounds, formulas, nomenclature of ionic compounds, covalent compounds, and polyatomic ions.	Chap. 2
08/24/18	F	3	Stoichiometry & Quantitative Chemistry, the mole concept, mass-%, chemical and empirical formulae	Chap. 3.1-3.2
08/27/18	M	3	Chemical equations, stoichiometry, balancing equations	Chap. 3.3
08/29/18	W	3	Chemical equivalents, calculating chem. quantities	Chap. 3.4
08/30/18	R	3	Review of chap. 3, stoichiometry, moles, limiting reactants	Chap. 3
08/31/18	F	4	Classes of Chemical Rxns, aqueous rxns, concentrations	Chap. 4.1-4.2
09/05/18	W	4	Precipitation rxns, acid/base rxns	Chap. 4.3-4.4
09/06/18	R	4	Acid/Base rxns cont'd, titrations, redox rxns	Chap. 4.4-4.5
09/07/18	F	4	Redox rxns cont'd, oxidation numbers, balancing redox rxns	Chap. 4.5-4.6
09/10/18	M	4	Reversibility and chem. equilibrium	Chap. 4.7
09/11/18	T		During-Term Exam #1, 8:20pm – 9:50pm.	Chap. 1 – 4
09/12/18	W	5	Gases, pressure, volume, ideal gas law, relationship between volume and amount of material	Chap. 5.1-5.3
09/13/18	R	5	Ideal gas law cont'd	Chap. 5.4
09/14/18	F	5	Kinetic gas theory, temperature, effusion, diffusion	Chap. 5.5
09/17/18	M	5	Real gases	Chap. 5.6
09/19/18	W	6	Thermochemistry, energy, energy conservation, heat and work	Chap. 6.1
09/20/18	R	6	Energy conservation cont'd.	Chap. 6.1
09/21/18	F	6	Enthalpy, calorimetry	Chap. 6.2-6.3
09/24/18	M	6	Stoichiometry of thermochemical rxns, Hess's Law, standard enthalpies	Chap. 6.4-6.6
09/26/18	W	16	Kinetics, rxn rates, measuring rates, rate laws, order of rate laws	Chap. 16.1-16.3
09/27/18	R	16	Integrated rate laws,	Chap. 16.3
09/28/18	F	16	Determination of rxn order, theories of chemical kinetics, collision theory, transition state theory, activation energy	Chap. 16.4-16.5
10/01/18	M	16	Rxn mechanisms, elementary rxns, molecularity, rate determining step approximation, steady state approximation, catalysis	Chap. 16.6-16.7
10/03/18	W	7	Nature of light, wave-particle duality, atomic spectra	Chap. 7.1-7.2
10/04/18	R	7	Bohr model, electron cloud, Heisenberg's uncertainty principle	Chap. 7.2-7.3
10/05/18	F	7	Quantum mechanical model of the atom, Schrödinger equation, orbitals, quantum numbers	Chap. 7.4
10/08/18	M	8	Electron configuration and periodicity, electron spin, Pauli exclusion principle	Chap. 8.1
10/09/18	T		During-term exam #2, 8:20pm – 9:50pm.	Chap. 5-7, 16
10/10/18	W	8	Penetration and shielding, Hund's Aufbau rules, electron configuration, transition elements	Chap. 8.2
10/11/18	R	8	Effective charge, ionization energies, periodic trends	Chap. 8.3
10/12/18	F	8	Periodic trends, cont'd	Chap. 8.3-8.4
10/15/18	M	9	Chemical Bonding, ionic bond, lattice energies, periodic trends	Chap. 9.1-9.2
10/17/18	W	9	Covalent bonds, bond enthalpy, bond order, bond lengths, reaction enthalpies	Chap. 9.3-9.4
10/18/18	R	9	Ionization energy, electronegativity, bond polarity	Chap. 9.5

10/19/18	F	9	Atomic periodicity in bonding, partial ionic character, metallic bond	Chap. 9.6
10/22/18	M	10	Lewis structures, orbitals, shapes of molecules	Chap. 10.1
10/24/18	W	10	VSEPR theory, orbitals	Chap. 10.2
10/25/18	R	10	Molecular shape and polarity	Chap. 10.3
10/26/18	F	11	Valence bond theory and orbital hybridization	Chap. 11.1
10/29/18	M	11	Orbital overlap, single vs. multiple bonds, bond rotation	Chap. 11.2
10/31/18	W	11	Molecular orbital theory, homo- and heterodiatomic molecules, σ -bonds, π -bonds, benzene, ozone	Chap. 11.3
11/01/18	R	11	Covalent bonds cont'd	Chap. 11.2-11.3
11/05/18	M	11	Review of molecular orbital theory	Chap. 11
11/05/18	M		During-Term Exam #3, 8:20pm – 9:50pm.	Chap. 8 – 11
11/07/18	W	12	Physical states and phase change, intermolecular forces, phase diagrams	Chap. 12.1-12.3
11/08/18	R	12	IMF's in liquids and solutions	Chap. 12.4
11/09/18	F	12	Water as a solvent	Chap. 12.5
11/14/18	W	12	The solid state of matter	Chap. 12.6
11/15/18	R	12	Advanced materials	Chap. 12.7
11/16/18	F	13	Types of solutions, IMF's in liquids and solutions	Chap. 13.1
11/19/18	M	13	IMF's in biological macromolecules, solvation	Chap. 13.2-13.3
11/26/18	M	13	Heat of solution, heat of hydration, entropy changes, solvation as an equilibrium process, concentration terms	Chap. 13.3-13.5
11/28/18	W	13	Colligative properties of solutions	Chap. 13.6
11/29/18	R	13	Colligative properties cont'd	Chap. 13.6
11/30/18	F	13	Colloids	Chap. 13.7
12/03/18	M		Review	Chap. 12 – 13
12/03/18	M		During-Term Exam #4, 8:20pm – 9:50pm.	Chap. 12 – 13
12/05/18	W		Review of the whole semester	
12/11/18	T		Cumulative Final Exam 10:00am – 12:00pm.	

FALL 2018

UNIVERSITY OF FLORIDA

SCHEDULE CARD

NAME: CHM2045 (Majors) Meeting Times and Office Hours

Univ. Address: Univ. Phone:

Per.	Hour	Monday	Bldg/Rm #	Tuesday	Bldg/Rm #	Wednesday	Bldg/Rm #	Thursday	Bldg/Rm #	Friday	Bldg/Rm #
1	7:25-8:15										
2	8:30-9:20										
3	9:35-10:25	CHM2045	JHH221	OH (Sam)	CLC*	CHM2045	JHH221	3F05	MAT119	CHM2045	JHH221
4	10:40-11:30			OH (Austin)	CLC*			3F11	LEI207		
5	11:45-12:35 p										
6	12:50-1:40	OH (Erik)	CLC*			OH (Erik)	CLC*			OH (Erik)	CLC*
7	1:55-2:45	OH (Jason)	CLC*			OH (Jason)	CLC*			OH (Sam)	CLC*
8	3:00-3:50	OH (Austin)	CLC*			OH (Dr. A)	JHH 202	OH (Dr. A)	JHH 202		
9	4:05-4:55			OH (Dr. A)	JHH 202						
10	5:10-6:00										
11	6:15-7:05										
E-1	7:20-8:10										
E-2	8:20-9:10										
E-3	9:20-10:10										

Comments: *Chemistry Learning Center, 1st floor of JHH, room #105.

Further Important Information:

1. **Overview and Goals:** CHM2045 is the first in a double course series introducing the student to General Chemistry. The goals of the course are to give an overview of basic chemistry, specifically atomic and molecular structure, stoichiometry, overview of chemical reactions, liquids and gases, thermochemistry, chemical kinetics, intermolecular forces, and colligative properties. The course is designed to prepare the student for subsequent work (General Chemistry 2 and Organic).
2. **Class Meeting Times:** The class meets in JHH221 MWF-3 period. Class discussion will start on time. Please be there a couple of minutes early. Small group discussions for the two sections meet during R-3 period in MAT 119 and during R-4 in LEI 207. Discussion will be led by one or more teaching assistants (TA).
3. **Majors Only:** The two sections, 3F05 and 3F11, of this course are only for Chemistry and Biochemistry majors. This allows for a smaller classroom environment more conducive to an interactive teaching style. It also allows for Chemistry/Biochemistry students to get to know each other sooner and interact within a small classroom environment. While the class will take the same exams as all other CHM2045 sections there are a few important differences to note. We will spend a little more time discussing topics that are important to Chemistry/Biochemistry majors down the road. That means that less time will be spent on drills. Hence, our Thursday discussion session will be used at about 50% to present new material while the other 50% is used to discuss homework problems. Homework sets will contain long-response problems that focus on understanding the material and students are encouraged to form study groups to work together on these problem sets. Worksheets with typical multiple-choice type GenChem questions will still be provided and are to be used on a voluntary self-study basis. Our class will use the 'Learning Catalytics' app for in-class participation while the other sections will use the 'tophat' app. Please do not purchase the tophat app unless you need it for some other course.
4. **General Chemistry Learning Objectives:** The course will provide instruction in the basic concepts, theories, and fundamental terms of chemistry. At the very core of chemistry is the concept of the atom, its structure, and bonding interactions with other atoms. Understanding the atom provides a conceptual foundation for the many aspects of 'macroscopic' chemistry. The focus of this course is understanding rather than memorization. Major scientific developments will be reviewed and their impacts on society, science, and the environment examined. Focus will be placed on the relevant processes that govern biological and physical systems. With what they learn students will be able to: (1) formulate empirically testable hypotheses relevant to the study of physical and life processes, (2) use logical reasoning skills through scientific criticism and argument, and (3) apply techniques of discovery and critical thinking to predict and evaluate outcomes of experiments. Upon successful completion of CHM2045 each student will:
 - have a working knowledge of the basic concepts, theories, and fundamental terms of Chemistry that are outlined under 1), and understand the relevant processes that govern chemical systems,
 - grasp the major scientific developments that have led to the current state-of-the-art in the field,
 - be able to assess impacts Chemistry has on society, science, and the environment,
 - be familiar with and capable of using the scientific method when discussing scientific facts as they relate to Chemistry,
 - know how to formulate empirically testable hypotheses derived from the study of physical and chemical processes,
 - use logical reasoning skills through scientific criticism and argument, and apply techniques of discovery and critical thinking to predict and evaluate outcomes of experiments.To achieve these objectives students are required to participate in all class activities, specifically:
 - Regular attendance of lectures in which the course material will be discussed and demonstrated is essential for success in the course. Lecture attendance requires active participation on the students' part. Large sections of class time will be spent in scientific dialog between teacher and students where we will practice the art of scientific reasoning.
 - One period of small group discussions is held each week in which students will discuss and apply the concepts learned in class under the guidance of a teaching assistant. The discussion sessions focus on homework problems and further explore difficult concepts that need additional explanation beyond the lectures. Students are expected to participate actively. This will not only give students helpful feedback on their own work but also train their logical reasoning skills through scientific criticism and argument.
 - Weekly graded homework assignments typically include between four and five conceptual and numerical problems that require the student to apply the learned concepts to specific examples. Problems are taken from different areas of experimental and theoretical chemistry including physical and life processes. Homework problems may also include reading material, typically a topically related original research article requiring the student to summarize and comment on in their own words.
 - 4 mid-term exams will be administered throughout the semester. These exams are the same exams that the other CHM2045 sections take as well. Due to time constraints and the fact that

thousands of students take these exams they can not be as detailed and time-intensive as the homework problems. Emphasis is placed on testing the students' reasoning skills and their understanding of the material rather than rote memorization of facts. On their exams, students will receive all pertinent equations.

- Participation points are available to students throughout the semester and can be earned both in the lectures and the discussion sessions by answering 'learning catalytics' questions (for more information see below item #10).
- Approximately 10 weekly office hours are offered by instructor and TAs, conveniently spread out over the whole week. Students are strongly encouraged to seek help and feedback on all concepts and problems encountered in class. While office hour attendance is completely voluntary, it is an important activity that will help solidify students' understanding of the material and make them successful in the course.

5. **Math Requirements:** Students are expected to have a solid grasp of pre-calculus algebra and trigonometry and need to have credit for MAC1147. Ideally, they should be co-registered for MAC2311 (Calculus 1). During lectures, concepts from Calculus may be briefly discussed but will not appear on exams. When derivatives or integrals are mentioned the focus is primarily on their graphical interpretation to aid understanding of chemical or physical concepts. When homework problems require their use, feel free to employ computational solvers such as Wolfram Alpha: <http://www.wolframalpha.com/>.
6. **ALEKS Prep:** Two percent of the course grade will be based on the Aleks prep work (due Monday, Sept. 10). For more info please see <https://www.chem.ufl.edu/undergraduate/aleks/>.

% ALEKS completion	0 – 69%	70 – 79%	80 – 89%	90 – 98%	99 – 100%
%-grade earned	0%	0.5%	1.0%	1.5%	2.0%

7. **Exam Policies:** Four during-term exams will be given (see schedule above). These exams will be evening exams. Exam duration will be approximately 1.5 hours. The final exam is cumulative. The Exam Room Assignments will be posted. You must use a non-graphing non-programmable scientific calculator on exams (with log, ln, root, and exponent (scientific notation) functions). Be sure to also bring pencils, section number, and your UF ID card. No notes, papers, cell phones or other electronic devices can be in view during exams. No makeup ("do over") progress exams will be given for any reason. If you must be absent for an exam due to a documented and approved academic or UF athletic conflict, bring the documentation to your instructor at least *one week prior* to the scheduled exam and an early conflict exam will be scheduled for you. If you are absent for an exam due to an unpredicted documented medical reason, you must contact the instructor as soon as possible and you have to get your excuse verified by the Dean of Student's Office. Your missed exam score will then be replaced by your pro-rated final exam score when calculating your final grade. More information regarding this policy can be found in the [General Chemistry Exam Absence Policy](https://www.chem.ufl.edu/wp-content/uploads/sites/38/2017/05/GenChemExamAbsencePolicy-05-05-2017.pdf) document found on the Chemistry Department web site: <https://www.chem.ufl.edu/wp-content/uploads/sites/38/2017/05/GenChemExamAbsencePolicy-05-05-2017.pdf>.

To alleviate the stress of potential issues that do not fall under officially-sanctioned absences, we have incorporated an "average/replace" policy (the lowest of the four progress exams will be replaced by the average of the four progress exams). This "average/replace" policy will help to minimize the impact of a single poor performance but it will not completely disappear. Any and all exam grade disputes or Scantron confirmations must be performed within two weeks of the scheduled exam date. Bubbling errors will not be negotiated, and a 5 point penalty will be applied for failure to bubble in a form code, UFID, or not taking the exam in the assigned room. University examination and reading day policies can be found at: <https://catalog.ufl.edu/UGRD/academic-regulations/examination-policies-reading-days/>

8. **Canvas:** Access your Canvas e-learning account by clicking on the 'Log-In to E-Learning' link on the web site, <http://lss.at.ufl.edu/> where you will have to supply your Gatorlink credentials to log in. Please, do this at your earliest convenience and make yourself familiar. Canvas will be primarily used by TAs and the instructor to communicate with the class. Please make sure to monitor the announcements on a regular basis. There may occasionally be assignments on Canvas that need to be completed before class. If you experience technical problems when using Canvas, e.g., during an online quiz, please contact the UFIT helpdesk (<http://helpdesk.ufl.edu/>, 352-392-4357 M-F from 8:00am till 5:00pm, email helpdesk@ufl.edu, or go to: <http://helpdesk.ufl.edu/e-learning-support/>).
9. **Homework (HW):** Ten HW assignments will be given over the course of the semester. They will normally be published on Canvas by Thursday afternoon and are due on the following Thursday at the beginning of discussion session. Late HW policy: HW is late if it is not delivered at the beginning of your discussion session to the TA in charge. Each day late will incur a 20% deduction of the total points value. Do your HW! By doing HW problems you will collect essential points toward your grade and will be better prepared to deal with problems on exams. HW problems come from many different sources, including the instructor's own personal list of problems. Since these will be the hardest problems you will encounter in CHM2045, you may form study groups with other students to work on them. However, simply copying someone else's work is plagiarism and

will be treated as such!

10. **Participation Grade:** Participation points (up to 5% of your total grade) will be earned through active participation in class with 'learning catalytics.' This is primarily done by using the app on your digital device (smartphone, tablet, notebook PC, *etc.*) to respond to questions asked by the instructor throughout the lectures (see further explanation below under #11).
11. **LearningCatalytics (LC):** In this course, we will use LC for your digital device to respond to the instructor's questions and earn valuable points toward your grade. You will need to purchase access and create a student account on <https://learningcatalytics.com/>. Follow instructions on that web site, or in the registration document on your Canvas account (click on Files → LC → Get_Started_Flyer_Learning_Catalytics.pdf) to activate your account and link it to our course, CHM2045. When registering make sure to use your full first and last name with correct spelling including capitalization. This is important to correctly link your LC gradebook entry to your course grades on Canvas. Do not register twice! If you have trouble logging in, get in touch with Pearson's help desk online at <https://support.pearson.com/getsupport/s/contactsupport>. The cost for LC is \$12 for the semester. You are required to bring at least one wifi-enabled digital device to class to use for this activity. If you don't have access to a digital device, please contact the instructor.
12. **Video Modules:** will be made available throughout the semester through links posted on Canvas. These modules are designed to aid in the explanation of concepts and will be used for instruction in addition to the classroom lectures. Typically, they should be watched before a topic is covered and announcements to that effect will go out to students ahead of time.
13. **Calculators:** You must have your own scientific calculator. Calculators may be used on homework and exams but may not be shared. You may **not** use graphing calculators or any calculators that are capable of communication on any exam. Simple inexpensive scientific calculators such as the TI-30 series or the Casio fx-260 are acceptable and sufficient for any problem encountered on exams.
14. **Class Attendance:** Regular attendance is essential for your success in this class. However, we will not do roll-calls. Repeated absence in class and discussion session will make it very difficult to earn full participation points. For further information on UF's attendance policies which are in effect for this course, see: <https://catalog.ufl.edu/ugrad/current/regulations/info/attendance.aspx>.
15. **Study Habits:** The course demands on average 10 – 12 hours/week of work outside of class. The class will not be taught 'by the book.' It is expected that you read the assigned pages from the textbook (or corresponding chapters in comparable textbooks) and watch assigned video modules in advance before coming to class. The instructor will build on this material and you are expected to be able to follow in-class discussion. The course demands a regular sustained effort throughout the semester. Most importantly, **do not allow yourself to fall behind!** The material builds up and you need to stay ahead of the game. If you find that you are not grasping essential material by reading the textbook and following in-class discussion, **seek help!** Visit your instructor's and/or TA's office hours, talk to other students in your class, compare notes, form a study group, consult other text books, go to the CLC (Chemistry Learning Center) in Joseph Hernandez Hall 105, *etc.*
16. **Study Groups:** It is highly encouraged to form study groups and meet with them on a weekly basis to discuss course material and to prepare for exams. In this course it is permissible that you work on HW assignments together with your study partners. However, you are responsible to fully understand your own worked-out HW submissions and may not just copy someone else's.
17. **Office Hours:** The instructor, two undergraduate TAs, and one graduate student TA offer a total of 10 office hours spread over the whole week. The detailed times and locations are listed on the first page of this syllabus and the table on the fourth page. This is time we set aside for you. Take advantage of it. Please note that the instructor and all TAs are available to help students in any of the two sections. You are not limited to only the TA assigned to your section. Moreover, the CLC (Chemistry Learning Center in JHH 105) is staffed with General Chemistry TAs (from other sections) throughout the week between periods 2 and 10 and you will be able to get help with your General Chemistry questions from all of them.
18. **Online Course Evaluation:** Students are expected to provide feedback on the quality of instruction in this course by completing online evaluations at <https://evaluations.ufl.edu>. Evaluations are typically open during the last two or three weeks of the semester. Announcements will be made to students about the specific times when they are open. Summary results of these assessments are available to students at <https://evaluations.ufl.edu/results/>.
19. **Students with Disabilities:** Students requiring special accommodations should register with the Dean of Students Office (<http://www.dso.ufl.edu/>, 352-392-1261) and the Disability Resource Center (DRC, <https://www.dso.ufl.edu/drc>, 352-392-8565, email: accessUF@dso.ufl.edu), and present documentation from that office to the instructor.
20. **Counseling Services:** The University of Florida provides counseling services for students, staff, and faculty. See <http://www.counseling.ufl.edu/cwc/>. If you or a friend are in distress, call (352) 392-1575 (available 24/7), email umatter@ufl.edu, or walk in for an emergency consultation during regular service hours (8:00am – 5:00pm) at the Radio Road Site, 3190 Radio Rd., or the Peabody Hall Site, on the 4th floor of Peabody Hall, adjacent to Criser Hall. For other hours or weekends, call the Alachua County Crisis Center, (352) 264-6789. For sexual assault recovery services call the Student Health Care Center at (352) 392-1161. For life-threatening emergencies always call 911.

21. **Emergency Numbers and Web Sites:**

- UFPD (UF Police Department): In case of emergency dial 911. The UF campus police non-emergency number is (352) 392-1111. Their web site: <http://www.police.ufl.edu/>,
- UF Emergency management: (352) 273-2100. <https://emergency.ufl.edu/>,
- Infirmary (student health center): (352) 392-1161, <http://shcc.ufl.edu/>.
- EH&S (Environmental Health & Safety): (352) 392-1591, <http://www.ehs.ufl.edu/>.

22. **Other Academic Resources:** UF provides several other resources for students, such as

- Library Support can be obtained here: <http://cms.uflib.ufl.edu/ask>, where you can find various ways to receive assistance with respect to using the libraries or finding resources.
- The Career Resource Center is located on level One in the Reitz Union, (352) 392-1601, and provides career assistance and counseling. Refer to <http://www.crc.ufl.edu/> for further info.
- The Teaching Center is located in Broward Hall, main phone (352) 392-2010 or appointment phone (352) 392-6420, and provides students with tutoring services and counseling regarding general study skills. Refer to <http://teachingcenter.ufl.edu/> for further info. It may also provide employment opportunities as tutors for well qualified students.
- The Writing Studio is located at 302, Tigert Hall, (352) 846-1138, and provides help with brainstorming, formatting, and writing papers, see: <https://writing.ufl.edu/writing-studio/>.
- The Ombuds Office is located at 31 Tigert Hall, (352) 392-1308, and provides students assistance in resolving problems and conflicts that arise in the course of interacting with the University of Florida. By considering problems in an unbiased way, the Ombuds works to achieve a fair resolution and works to protect the rights of all parties involved. For further information go to <http://www.ombuds.ufl.edu/> or refer to the official complaints policy here: https://www.dso.ufl.edu/documents/UF_Complaints_policy.pdf.

23. **Cell Phone Etiquette:** Please put all cell phones or other electronic devices on “**silent mode**” during all class and discussion periods. Please do not leave the classroom during lecture to make a phone call. Use your cellphone only for ‘learning catalytics’ activities while class is in session. Thank you!

24. **Facebook Page:** One of the undergraduate TAs has set up a closed facebook group for the class here: <https://www.facebook.com/groups/457011178143266/>. Participation through reading and posting is voluntary but will enhance the class community. Please note that this page is not monitored by the instructor. It is administered by the TAs.

25. **Honor Code:** This class will operate under the policies of the student honor code which can be found at: <https://www.dso.ufl.edu/sccr/process/student-conduct-honor-code/>. The students, instructor, and TAs are honor-bound to comply with the Honors Pledge: **We, the members of the University of Florida community, pledge to hold ourselves and our peers to the highest standards of honesty and integrity.** You are expected to exhibit behavior consistent with this commitment to the UF academic community, and on all work submitted for credit at the University of Florida, the following pledge is either required or implied: *"On my honor, I have neither given nor received unauthorized aid in doing this assignment."* It is assumed that you will complete all work independently in each course unless the instructor provides explicit permission for you to collaborate on course tasks. Furthermore, as part of your obligation to uphold the Honor Code, you should report any condition that facilitates academic misconduct to appropriate personnel. It is your individual responsibility to know and comply with all university policies and procedures regarding academic integrity and the Student Honor Code. Violations of the Honor Code at the University of Florida will not be tolerated. Violations will be reported to the Dean of Students Office for consideration of disciplinary action. For more information regarding the Student Honor Code, please see: <https://www.dso.ufl.edu/sccr/process/student-conduct-honor-code/>.

26. **Disclaimer:** This syllabus represents my current plans and objectives. If those need to change as the semester progresses, then the changes will be communicated to the class clearly.

If you have further questions, please contact me. Have a great semester!

Sincerely,

Alexander Angerhofer
(Dr. A)