Syllabus CHM 3120L ANALYTICAL CHEMISTRY LABORATORY Spring, 2016

Faculty Instructor: Dr. Ben Smith, Keene-Flint 264, 392-0256, bwsmith@ufl.edu, Office Hours: M, W, F period 3

You are welcome to visit me at any time in my office.

Teaching Assistants: Imran Iftikar <u>imif6145@ufl.edu</u>

Emily Gill emilygill2014@ufl.edu
Wenbo Peng wenbopeng@ufl.edu
Dan Mark djm17@ufl.edu
Long Li lilong@ufl.edu

Undergraduate Teaching Assistants:

Abel Abraham abelabra@ufl.edu Alec MacKinnon alec118@ufl.edu msciturro@ufl.edu Monica Sciturro Melissa Andrews mandrews24@ufl.edu afriedman4@ufl.edu Austin Freidman velloww7@ufl.edu Alaina Giacobbe Meghan White mookers@ufl.edu snagesh@ufl.edu Surya Nagesh Ciara.sanon@ufl.edu Ciara Sanon Alexandria Shrewsbury alishrews@ufl.edu Jacob Dautel jacobadautel@ufl.edu Madeleine Morin m.morin@ufl.edu Marc Nebb mnebb1@ufl.edu

Course Website: Canvas; Please visit the website regularly for announcements and resources.

Videos available at: http://www.chem.ufl.edu/ugrad/labanalytical.shtml

Required Materials

Laboratory Manual: No lab manual is required. All materials will be posted on the Canvas e-learning site, under

Resources.

Laboratory Notebook: Any sensible laboratory notebook, to be used only for this lab, is suitable. You will turn in either

carbon copies, or scans or Xerox copies of your notes, retaining the original notebook for your own

use. Please be sure that what you submit is legible and clear.

Laboratory Attire: The Essentials: Long, loose-fitting pants, full shirt, shoes which cover the feet, departmentally-

approved safety glasses, tie-back for long hair.

Course Objectives

CHM 3120L is an introductory laboratory course in Analytical Chemistry. By the end of the semester, students are expected to demonstrate:

- proper laboratory techniques for quantitative chemical measurements including accuracy on unknowns
- knowledge of a select group of analytical methods
- competence in data analysis and preparation of basic laboratory reports

.

Grading

Your grade will be determined by the accuracy of your results, the quality of your reports, the quality of your laboratory notes, your competence in essential laboratory manipulations, and your performance on written quizzes.

```
Accuracy 5 @ 70 points 350
Reports and Notes 6 @ 70 points 350
Practical Exams 3 @ 40 points 120
Written Quizzes 4 @ 45 points 180
1000 total
```

The following grading scale will be used:

```
A\ (88.0-100\%),\ A-\ (86.0-87.9\%),\ B+\ (81.5-85.9\%),\ B\ (78.5-81.4\%),\ B-\ (74.5-78.4\%),\ C+\ (71.5-74.4\%),\ C\ (67.0-71.4\%),\ C-\ (64.5-66.9\%),\ D+\ (60.0-64.4\%),\ D\ (57.0-59.9\%),\ D-\ (53.0-56.9\%),\ E\ (<53.0\%).
```

Notes:

- Prior to the first lab, visit the e-learning site and review Preliminary Handouts 1-5: laboratory safety, basic lab rules, laboratory notebook, laboratory reports and fundamental techniques. Also read the handout for Experiment #1.
- 2) A minimum of 30 out of 70 accuracy points will be given if the experiment is performed, the results are calculated correctly and deadlines are met.
- 3) For each of the six experiments you will write concise formal laboratory reports. Reports are due at the beginning of your laboratory period during the week specified. The laboratory experimental guidelines will contain questions for each experiment. These are designed to help you prepare for the written quizzes. Written answers are not required as part of the reports.
- 4) A 10 point penalty will be assessed each time a result or report is submitted late. The maximum permissible late time is one week.
- Each student is expected to pass laboratory practical exams on three essential analytical skills (use of the analytical balance/weighing by difference, quantitative transfer/use of a volumetric flask and use of a pipets). The tests will be given by the TA during the regular laboratory period at times mutually acceptable to both the student and the TA.
- 6) Four written quizzes will be given on the dates specified on the schedule. Study material will be posted one week in advance. You will be allowed to see your graded written quiz, but it must be returned to the TA before leaving lab.
- 7) Attendance is required at all scheduled laboratory periods, unless you are informed otherwise by your TA.
- 8) Once an unknown result has been submitted, no repeat work on that unknown is allowed.
- 9) Students are expected to obey the University of Florida Honor Code, detailed at https://www.dso.ufl.edu/sccr/process/student-conduct-honor-code/. Violations will be reported to the Office of Student Judicial Affairs.
- Make-ups will be granted only when justified. If you know ahead that you will have to miss lab, notify your TA and Dr. Smith in advance. If you are sick and cannot reach anyone before lab, you will have to present written evidence of the illness.
- 11) If you are involved in a laboratory accident, you must go to the infirmary for treatment.
- 12) Students requesting classroom accommodation must first register with the Dean of Students Office. The Dean of Students Office will provide documentation to the student who must then provide this documentation to the instructor when requesting accommodation.
- 13) Lab preparatory videos are available at: http://www.chem.ufl.edu/ugrad/labanalytical.shtml

Laboratory Schedule

Note: Because of holidays, different sections may be working on different experiments during any given week. Regardless,

each section will follow the sequence of activities given below.

| Dates | Preparation | Lab Work | Quizzes | Results Due |
|------------------|-----------------------|---|------------------------|----------------------|
| Begin January 11 | Read Handouts 1-6 | Check in | | |
| Week 1 | Read Experiment 1 | Experiment 1 | | |
| | Watch video: Lab | Balance use | | |
| | Techniques | Pipet use/calibration | | |
| Week 2 | Read Handout 7 | Begin Soda Ash Titrations | | #1 Calibrations |
| | Read Experiment 2 | HCl/NaOH titrations | | (report and lab |
| | | KHP/NaOH titrations | | notes) |
| Week 3 | | Finish Soda Ash Titrations | Quiz 1 and | |
| | | | Deadline for Weighing | |
| | | | Practical | |
| Week 4 | Watch ascorbic acid | Prep KIO ₃ and Na ₂ S ₂ O ₃ | | #2 Soda Ash (report |
| | video | Standardize Na ₂ S ₂ O ₃ | | and lab notes) |
| | Read Experiment 3 | | | |
| | Look over handout for | Standardize Na ₂ S ₂ O ₃ | | |
| | Quiz 1 | Ascorbic acid titrations | | |
| Week 5 | | Finish Ascorbic acid | Quiz 2 | |
| | | | Deadline for Pipetting | |
| | | | Practical | |
| Week 6 | Read Experiment 4 | Spectrophotometric Fe | | #3 Ascorbic Acid: |
| | Watch Spec Fe video | • | | Report and lab notes |
| | • | | | • |
| Week 7 | | Finish Fe | | |
| Week 8 | Watch ISE video | Ion Selective Electrode | Quiz 3 and deadline | #4 Spec Fe: report |
| | Read Experiment 5 | | for volumetric flask | and lab notes |
| | • | | practical | |
| Week 9 | Read Experiment 6 | Kinetics | | #5 ISE: report and |
| | - | | | lab notes |
| Week 10 | | Finish kinetics | | |
| Week 11 | | Check Out | Quiz 4 | #6 Kinetics: report |
| | | | | and lab notes |